

Complete the table to show the difference between metals and non-metals

| Property                           | Metals | Non Metals |
|------------------------------------|--------|------------|
| Conduction of Heat/Electricity     |        |            |
| Ability to bend (Malleable or not) |        |            |
| Appearance                         |        |            |
| Sound when hit                     |        |            |
| Density                            |        |            |
| Magnetism                          |        |            |

In the space below draw an atom and label it with the keywords



Complete the table to show the properties of the Alkali Metals (Group 1)

## Atomic Structure and Periodic Table

Why are Group 0 atoms unreactive?

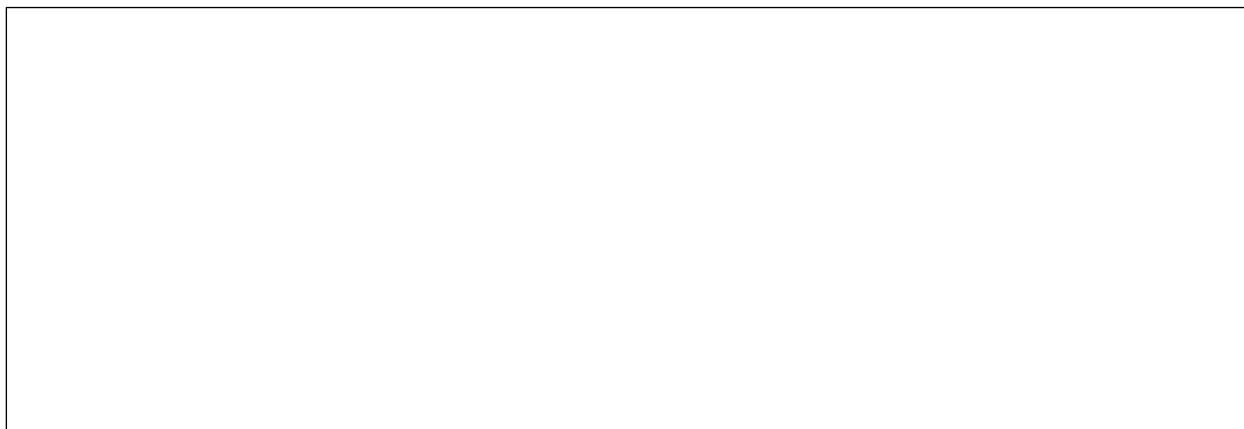
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Complete the table to show the properties of the Halogens (Group 7)

|                                    | Fluorine | Chlorine | Bromine | Iodine |
|------------------------------------|----------|----------|---------|--------|
| Electron Structure                 |          |          |         |        |
| State at room temperature          |          |          |         |        |
| Key information about the halogens |          |          |         |        |
| Reactivity                         |          |          |         |        |
| Products when reacted with metals  |          |          |         |        |

|                                 | Lithium | Sodium | Potassium |
|---------------------------------|---------|--------|-----------|
| Electronic Structure            |         |        |           |
| Melting point/ °C               |         |        |           |
| Flame Test Colour               |         |        |           |
| Reaction with water             |         |        |           |
| Products of reaction with water |         |        |           |

Draw a timeline to describe how the periodic table was developed through time. Include the scientists and their theories.

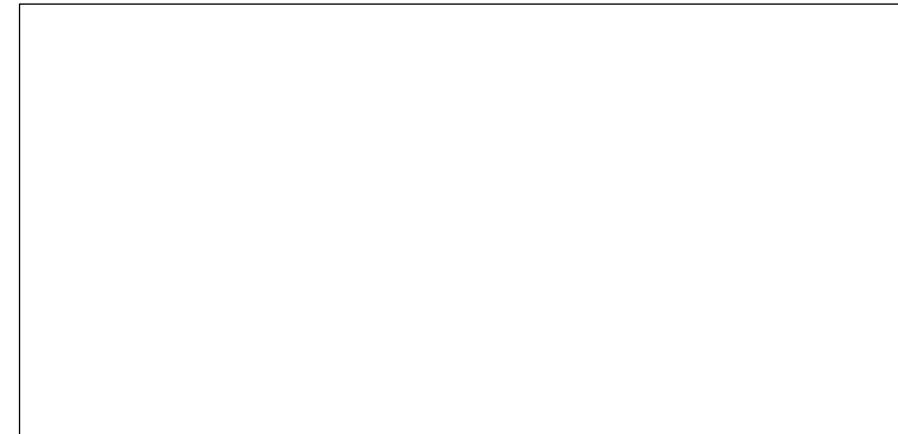


# Atomic Structure and Periodic Table

List 4 properties of transition metals

Shade the periodic table below to show where the transition metals are found

| Group →     | 1        | 2        | 3        | 4         | 5         | 6         | 7         | 8         | 9         | 10        | 11        | 12        | 13         | 14         | 15         | 16         | 17         | 18         |
|-------------|----------|----------|----------|-----------|-----------|-----------|-----------|-----------|-----------|-----------|-----------|-----------|------------|------------|------------|------------|------------|------------|
| ↓ Period    |          |          |          |           |           |           |           |           |           |           |           |           |            |            |            |            |            |            |
| 1           | 1<br>H   |          |          |           |           |           |           |           |           |           |           |           |            |            |            |            | 2<br>He    |            |
| 2           | 3<br>Li  | 4<br>Be  |          |           |           |           |           |           |           |           |           |           |            |            |            |            |            |            |
| 3           | 11<br>Na | 12<br>Mg |          |           |           |           |           |           |           |           |           |           |            |            |            |            |            |            |
| 4           | 19<br>K  | 20<br>Ca | 21<br>Sc | 22<br>Ti  | 23<br>V   | 24<br>Cr  | 25<br>Mn  | 26<br>Fe  | 27<br>Co  | 28<br>Ni  | 29<br>Cu  | 30<br>Zn  | 31<br>Ga   | 32<br>Ge   | 33<br>As   | 34<br>Se   | 35<br>Br   | 36<br>Kr   |
| 5           | 37<br>Rb | 38<br>Sr | 39<br>Y  | 40<br>Zr  | 41<br>Nb  | 42<br>Mo  | 43<br>Tc  | 44<br>Ru  | 45<br>Rh  | 46<br>Pd  | 47<br>Ag  | 48<br>Cd  | 49<br>In   | 50<br>Sn   | 51<br>Sb   | 52<br>Te   | 53<br>I    | 54<br>Xe   |
| 6           | 55<br>Cs | 56<br>Ba |          | 72<br>Hf  | 73<br>Ta  | 74<br>W   | 75<br>Re  | 76<br>Os  | 77<br>Ir  | 78<br>Pt  | 79<br>Au  | 80<br>Hg  | 81<br>Tl   | 82<br>Pb   | 83<br>Bi   | 84<br>Po   | 85<br>At   | 86<br>Rn   |
| 7           | 87<br>Fr | 88<br>Ra |          | 104<br>Rf | 105<br>Db | 106<br>Sg | 107<br>Bh | 108<br>Hs | 109<br>Mt | 110<br>Ds | 111<br>Rg | 112<br>Cn | 113<br>Uut | 114<br>Uup | 115<br>Uuh | 116<br>Uus | 117<br>Uuo | 118<br>Uuo |
| Lanthanides |          |          |          |           |           |           |           |           |           |           |           |           |            |            |            |            |            |            |
| Actinides   |          |          |          |           |           |           |           |           |           |           |           |           |            |            |            |            |            |            |
|             | 57<br>La | 58<br>Ce | 59<br>Pr | 60<br>Nd  | 61<br>Pm  | 62<br>Sm  | 63<br>Eu  | 64<br>Gd  | 65<br>Tb  | 66<br>Dy  | 67<br>Ho  | 68<br>Er  | 69<br>Tm   | 70<br>Yb   | 71<br>Lu   |            |            |            |
|             | 89<br>Ac | 90<br>Th | 91<br>Pa | 92<br>U   | 93<br>Np  | 94<br>Pu  | 95<br>Am  | 96<br>Cm  | 97<br>Bk  | 98<br>Cf  | 99<br>Es  | 100<br>Fm | 101<br>Md  | 102<br>No  | 103<br>Lr  |            |            |            |



Compare the properties of transition metals with those of Group 1 metals

