Identify how many atoms are in each compound / molecule and calculate the Relative Formula Mass (Mr).

Formula	No. of Atoms	Mr
CO_2		
O_2		
MgO		
$CaCl_2$		
$CuSO_4$		
$Mg(NO_3)_2$		

Balance the equations below

$$CuO + C \longrightarrow Cu + CO_2$$

$$Mg + O_2 \longrightarrow MgO$$

$$Ca + HCl \longrightarrow CaCl_2 + H_2$$

$$Li + H20 \longrightarrow LiOH + H_2$$

$$Al_3O_2 + H_2SO_4 \longrightarrow Al_2(SO_4)_3 + H_2O$$

Describe what the conservation of energy is

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Quantitative Chemistry

Work out the % mass of the following:
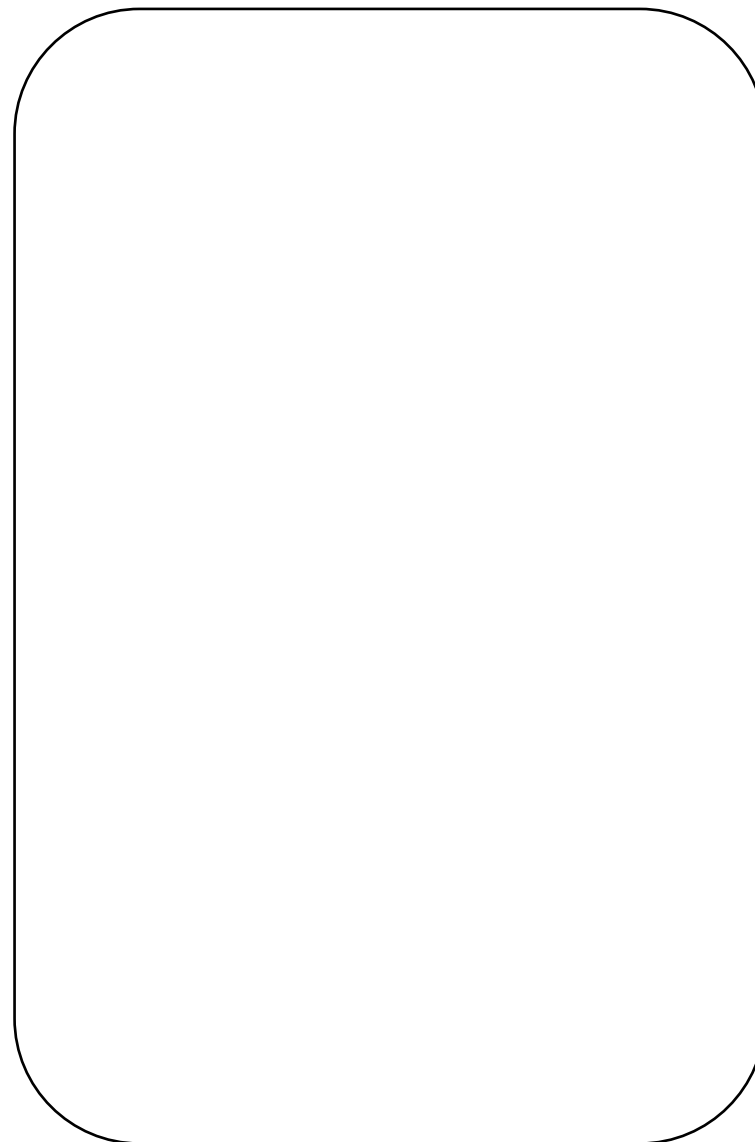
% of Mg in MgO

% of Li in LiOH

% of H in CH_4

% of O in CO_2

In the space below describe a practical with a diagram that can be done to show the conservation of mass principle.