

C1 REVISION - CHAPTER 4 - CHEMICAL CHANGES - METALS

1) Complete the reactivity series below:

Potassium
Lithium
Calcium
Aluminium
Iron
Tin
Lead
Silver
Gold



_____ reactive

2) Explain how the metals in the reactivity series react with water.

3) Explain how the metals in the reactivity series react with acid.

4) What is a displacement reaction?

5) Predict which metals will displace zinc from a solution.

6) Complete the following equation:

Iron + Lead nitrate \rightarrow _____ + _____

7) Add hydrogen and carbon to the reactivity series on the left.

Describe how metals are extracted by carbon, using equations.

Describe how metals are extracted by hydrogen, using equations.

Complete the anagram below:-

o
i
l
r
i
g

Explain why carbon can reduce zinc oxide but magnesium oxide cannot